

INDEX

- 1. Details of the Students**
- 2. Syllabus**
- 3. Pre Bridge course Test**
- 4. Pre Bridge course Test Analysis**
- 5. Day wise Schedule**
- 6. Post Bridge course Test**
- 7. Post Bridge course Test Analysis**

Syllabus

Organic Chemistry

- Fundamental particles of an atom
- Bohr's atomic theory
- Quantum Numbers
- Basic rules for electronic Configuration
- Atomic number-Electronic configuration of Elements
- Valency of carbon
- Types of hybridization in carbon compounds
- Pi bond formation – bond polarisation
- Inductive effect
- Mesomeric effect
- Hyper conjugation effect
- Electronic configuration of Elements
- Sigma and Pi bond formation
- Valency bond theory
- Hybridisation of orbitals with examples




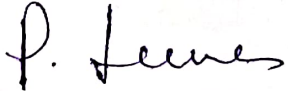

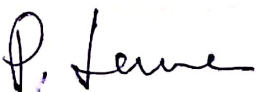

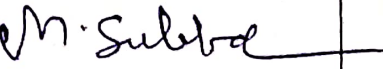
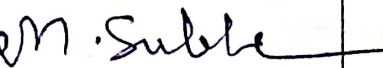

Physical Chemistry

- Definition of Lattice point, Space lattice, Unit Cell
- Bragg's Law
- Defects in Crystals
- Joule Thomson effect
- Liquid Crystals
- Nernst Distribution Law
- Common Ion Effect
- Solubility Product
- Colligative properties

Inorganic Chemistry

- Periodic Table
- Diborane Structure
- Oxidation states
- Magnetic Properties
- Lanthanide Contraction

Sl.No	Date	
1.	31.10.2022	Boh Nur Cor
2	01.11.2022	Va Hy Po
3	02.11.2022	In Et
4	03.11.2022	V C
5	04.11.2022	I S I
6	05.11.2022	
7	07.11.2022	
8	08.11.2022	
9	09.11.2022	
10	10.11.2022	
11	11.11.2022	

Date	Name of the Topic Covered	Name of the Lecturer	Signature
31.10.2022	Bohr's Model, Quantum Numbers, Electronic Configuration	Dr.K.Jhansi Lakshmi	
01.11.2022	Valency of Carbon, Hybridization, Bond Polarization	Dr.K.Jhansi Lakshmi	
02.11.2022	Inductive Effect, Mesomeric Effect, Hyper conjugation	Dr.K.Jhansi Lakshmi	
03.11.2022	VB Theory, Hybridisation of Orbitals with examples.	P.Leena	
04.11.2022	Definition of Lattice Point, Space Lattice, Unit cell, Bragg's Law, Crystal Defects	P.Leena	
05.11.2022	Joule Thomson Effect, Liquid Crystals, Nernst Distribution Law	P.Leena	
07.11.2022	Common Ion Effect, Solubility Product, Colligative properties	M.Subbalakshmi	
08.11.2022	Periodic table, Diborane structure	M.Subbalakshmi	
09.11.2022	Oxidation States, Magnetic Properties	M.Subbalakshmi	
10.11.2022	Lanthanide contraction and their consequences	M.Subbalakshmi	
11.11.2022	Test conducted	K. Jhansi lakshmi P.Leena M.Subbalakshmi	

Students Details

Sl.No	Name of the Student	Group
1	V.Vijaya Lakshmi	MPC
2	K.Veera Veni	MPC
3	Ch.Madhuri	MPC
4	A.Lakshmi	CBMB
5	P.Bhavani	CBMB
6	J.Suji	CBZ
7	R.R.Venkatalakshmi	CBZ
8	K.Sharmila Ganga	CBZ
9	S.Meghana Sriveni	CBZ
10	D.Ganga Bhavani	MPC
11	K.Bhavani	MPC
12	S.Sammakka	MPC
13	T.Anantha	MPC
14	K.Anusha	CBHT
15	R.D.Kumari	CZAqT
16	K.Bala Rajini	CZAqT
17	V.Muneeswari	CZAqT
18	V.Sudha Rani	CZAqT
19	Ch.Anitha	CZAqT
20	M.Rani	CBMB
21	P.Kusuma	CBMB

22	P.Susma	CBMB
23	B.P.Pushpa	CBMB
24	R.Madhu	CBMB
25	I.Srilakshmi	CBHT
26	S.Raja Kumari	CBHT
27	K.Sai Kumari	MPC
28	A.D.Mahalakshmi	MPC
29	M.Satya Spandana	MPC
30	B.Lovatalli	MPC
31	R.Sunitha	MPC
32	P.Krishna Veni	CBMB
33	K.Veera Veni	CBHT
34	B.Durga Bhavani	CBZ
35	A.Akhila	CBZ
36	K.Saranya	CBHT
37	T.Asha Jyothi	CBHT
38	S.Prema Vani	CBHT
39	G.Akshaya	CBHT
40	P.Durga Bhavani	CBZ

A.S.D GOVERNMENT DEGREE COLLEGE FOR WOMEN(A),
KAKINADA

DEPARTMENT OF CHEMISTRY
BRIDGE COURSE QUESTIONNAIRE

2022-2023

- Who introduced the electron ?
A) J.J.Thompson B) Rutherford C) Chadwick D) Newton
- Which Quantum number determines the shape of the orbital.
A) Principal quantum number B) Angular Quantum Number
C) Magnetic Quantum Number D) None of these
- Electronic configuration of inert gas is
A) $1S^2 2S^2$ B) $1S^2 2S^2 2P^3$ C) $1S^2 2S^2 2P^6$ D) $1S^2 2S^2 2P$
- Which group is known as Alkali metal family ?
A) VII A B) IA C) IIA D) VIIIA
- What is the valency of carbon atom ?
A) 2 B) 3 C) 4 D) None of these
- Highest electronegative element
A) Cs B) F C) Cl D) Br
- what is the Hybridation in H_2O molecule?
A) SP^3 B) SP C) SP D) SP^3d
- Which among the Following is a non metal
A) Potassium B) Chlorine C) Silicone D) Sodium
- The Maximum number of electrons in a sub shell is given by
A) $2l+1$ B) $2(2l+1)$ C) $3n+1$ D) $2n^2$
- Which of the following is not a Crystalline solid ?

- A) Kcl B) Cscl C) Glass D) Rhombic Sulphur
11. Which substance will conduct the current in the solid state ?
 A) Diomond B) Graphite C) Iodine D) Sodium
12. Which Defect causes in the density of the crystal ?Which
 A) Frenkel B) Schotty C) F centre D) Interstial
13. Which of the following has no units ?
 A)Morality B)Normality C)molality D)Mole Fraction
14. Which of the following is a colligative property
 A) Boiling Point B)Osmotic Pressure
 C)Vapour pressure D)Freezing Point
15. Lanthanoids and Actinoids together belong to
 A)S - Block B)P - Block C)D- Block D)F – Block
16. Electronic Configuration of Chromium.
 A) $(Ar^{18})3d^54s^1$ B) $(Ar^{18})3d^44s^2$ C) $(Ar^{18})3d^94s^2$ D) $(Ar^{18})3d^{10}4s^2$
17. Hybridisation of Carbon in Acetylene
 A) SP^3 B) SP^2 C) SP D) SP^3d
18. Oxidation state of Manganese in $KMnO_4$
 A) +2 B) +7 C) +6 D) 0
19. Bond length of Carbon - Carbon double bond.
 A) 1.54 \AA B) 1.34 \AA C) 1.30 \AA D) 1.20 \AA
20. Which one is not a inert gas.
 A) He B) Pt C) Ar D) Kr

KEY

1.A	4.B	7.A	10.C	13.D	16.A	19.B
2.B	5.C	8.B	11.D	14.B	17.C	20.B
3.C	6.B	9.B	12.B	15.D	18.B	

Pre and Post Bridge Course Test Marks

S.No	Name of the Student	Pre Bridge Course Test	Post Bridge Course Test
1	V.Vijaya Lakshmi	15	16
2	K.Veera Veni	15	18
3	Ch.Madhuri	13	17
4	A.Lakshmi	13	15
5	P.Bhavani	14	17
6	J.Suji	13	18
7	R.R.Venkatalakshmi	13	19
8	K.Sharmila Ganga	13	20
9	S.Meghana Sriveni	14	17
10	D.Ganga Bhavani	15	16
11	K.Bhavani	10	16
12	S.Sammakka	11	17
13	T.Anantha	10	17
14	K.Anusha	15	18
15	R.D.Kumari	11	19
16	K.Bala Rajini	10	20
17	V.Muneeswari	13	19
18	V.Sudha Rani	11	20
19	Ch.Anitha	10	16
20	M.Rani	15	18
21	P.Kusuma	15	17

22	P.Susma	11	16
23	B.P.Pushpa	10	18
24	R.Madhu	9	17
25	I.Srilakshmi	14	17
26	S.Raja Kumari	8	18
27	K.Sai Kumari	10	16
28	A.D.Mahalakshmi	11	17
29	M.Satya Spandana	15	18
30	B.Lovatalli	14	20
31	R.Sunitha	13	19
32	P.Krishna Veni	12	19
33	K.Veera Veni	10	20
34	B.Durga Bhavani	11	17
35	A.Akhila	09	16
36	K.Saranya	08	15
37	T.Asha Jyothi	07	18
38	S.Prema Vani	06	16
39	G.Akshaya	10	17
40	P.Durga Bhavani	10	18

V. M. R.